



# DRIVE Marine Services

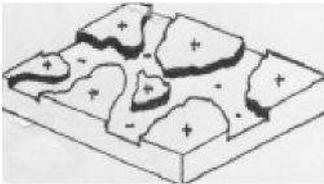
## FERONITE HOW IT WORKS

### HISTORY

It has been known for many years that vegetable tannins possess useful anti-corrosive properties. Old carpenters used to pack their tools in wood shavings to prevent them rusting. Everyone has seen how a nail driven into green timber develops a black stain on its surface and in the timber around it - this is the reaction product of iron oxide and the tannins in the wood, and it slows down the rusting process of the nail.

Chemical research into tannin technology was taking place in Europe during the 1970's, and Feronite is a result of that research.

### HOW FERONITE STOPS RUST

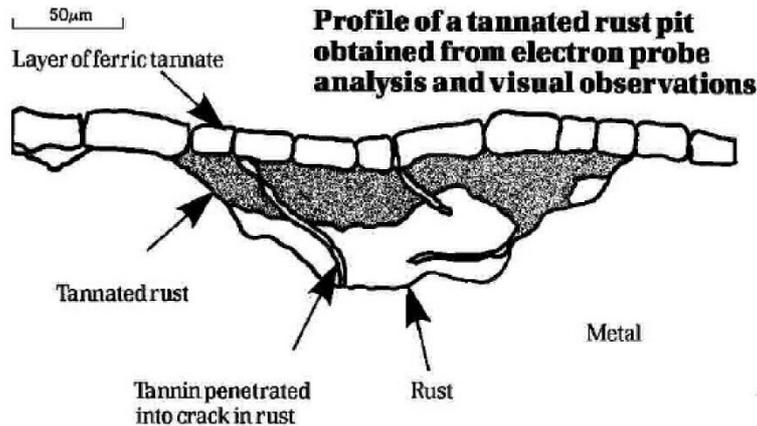


An iron surface is made up of microscopic positive (anodic) and negative (cathodic) areas. In the presence of moisture, they react together and cause the surface to form a mixture of iron oxides. The anodic area is the one where metal actually reacts or dissolves to form the oxide, but, as the chemical nature of the surface continually changes, the cathodic and anodic areas tend to move around so that corrosion takes place over all of the surface.

Feronite reacts with rust to produce a different form of iron oxide together with iron tannate. This layer is both insoluble and passive to further corrosion.

The insolubility inhibits water penetration, an important factor in the rusting process. The passive nature of the iron tannate stops the movement of the anodic and cathodic areas, which also helps stop the spread of the rust.

Scientific analysis of a rusted surface shows a picture of a reduced oxide (ferrous oxide) at the steel interface covered with tannated iron oxides and a final reaction of ferric tannate on the outer surface. The whole surface is more strongly bonded than the original ferric oxide rust layer, and the scaly or 'reticulated' ferric tannate surface acts as an excellent key for follow on paint systems.



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